

Electron Configuration – Worksheet 2

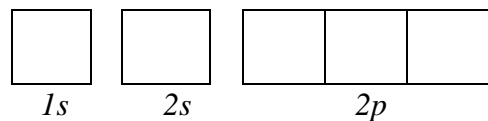
1. Explain Aufbau's principle, Hund's rule, and the Pauli exclusion principle. Provide one example for each principle and rule.

2. For each of the following atoms, write the full electron configuration and draw the corresponding orbital diagram.

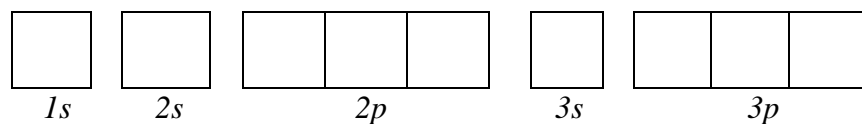
Atom	# of Electrons	Electron Configuration and Orbital Diagram
Mg		
Ar		
P		
Ge		
Ca		
Cl		

3. Write the full electron configuration, short-hand electron configuration, and fill in the orbital diagrams, for the following elements.

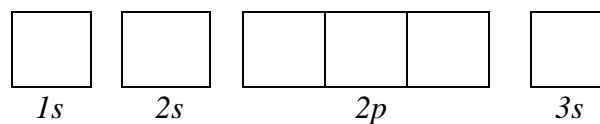
1. Nitrogen



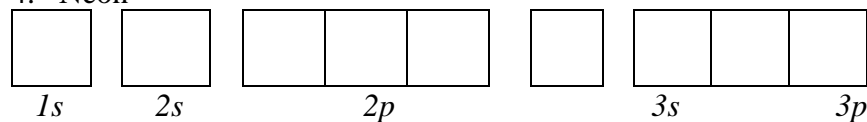
2. Sulfur



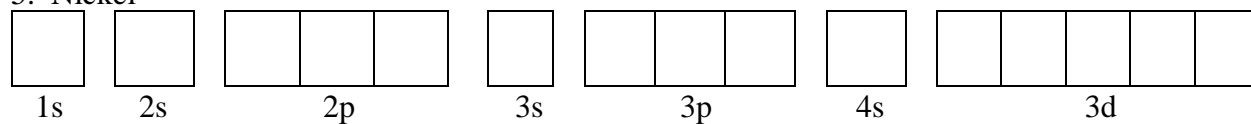
3. Sodium



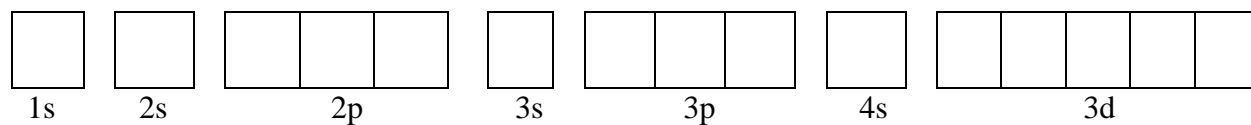
4. Neon



5. Nickel



6) Iron



7) Zinc

